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3-1 Calculate the specific volume of an air-vapor mixture in cubic meters per kilogram of dry air when the following conditions prevail: $t = 30\text{ C}$, $W = 0.015\text{ kg/kg}$, and $p_t = 90\text{ kPa}$. Solution: Equation 3-4. $t_s = 30\text{ C} + 273 = 303\text{ K}$ $R_a = 287\text{ J/kg.K}$ $p_t = 90\text{ kPa} = 90,000\text{ Pa}$ Equation 3-2 $t_s = 30\text{ C}$ $t_1 = -10\text{ C}$ $\frac{p_1}{p_2} = \left(\frac{T_1}{T_2}\right)^{\frac{1}{\gamma}}$ $\frac{90}{0.622p} = \left(\frac{263}{303}\right)^{\frac{1}{1.4}}$ $0.622p = \frac{90}{\left(\frac{263}{303}\right)^{\frac{1}{1.4}}}$ $0.622p = 1.35$ $p = 2.17\text{ kPa}$

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Solution: $q = mc_p(t_2 - t_1)$ $m = 2.5\text{ kg/s}$ $c_p = 1.0\text{ kJ/kg.K}$ $t_2 = 30\text{ C}$ $t_1 = -10\text{ C}$ CHAPTER 2 - THERMAL PRINCIPLES Page 2

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of 5 Then, $q = (2.5)(1.0)(30 + 10)q = 100 \text{ kw} - - - \text{Ans. 2-3.}$

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