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first law as  $w = q - \Delta U$

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; thus  $q = DH$ ;  $w = D$   
HPVL; 4. Isothermal  
Process Because  $U$  is  
a function only of  $T$  for  
an ideal gas,  $DU =$   
 $DH = 0$  for an  
isothermal process.  
These results also  
follow from the  
general results by  
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Problem 1.1\* The plot  
of  $V = V(P, T)$  for a  
gas is shown in Fig.  
1.1. Determine. the

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expressions of the two second derivatives of the volume of this plot. (note: the principle curvatures of the surface are proportional to these second derivatives).

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Work is found the the  
first law as  $w = q -$   
 $DU$ ; thus  $q = DH$ ;  $w =$   
 $DHPVL$ ; 4.

Isothermal Process

Because  $U$  is a  
function only of  $T$  for  
an ideal gas,  $DU =$   
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general results by  
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The isothermal  
expansion is

conducted at 300 K.

3.4  $\Delta H = 42750 \text{ J}$ ,  $\Delta S$

$= 59.7 \text{ J/K}$  3.5 The

final temperature is

323.32 K, which is

greater than 323 K

because the heat

capacity increases

with increasing

temperature.

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± Physical Constants  
Needed for Problems

ü Heat Capacities The  
generic heat capacity

$$C_p = a + b T \\ \text{ÅÅÅÅÅÅÅÅÅÅÅ} 10^3 + \\ c 10^5$$

ÅÅÅÅÅÅÅÅÅÅÅÅÅÅÅÅ

Å T<sup>2</sup> ; The heat  
capacities of various  
elements and  
compounds are

$$C_{p, Ag} = C_p \hat{e}. 8a \emptyset$$

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21.30,  $b \neq 8.54$ ,  $c \neq$

$1.51$ ;  $C_p A_{gl} = C_p \hat{e}$ .

$8a \neq 30.50$ ,  $b \neq 0$ ,  $c$

$\neq 0$ ;  $C_p A_l = C_p +$

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