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This reaction follows the equation: $N_2 + 3H_2 \rightarrow 2NH_3$ (1)
 $H_f^0 = 92 \text{ kJ mol}^{-1}$ $S_f^0 = 192 \text{ J mol}^{-1} \text{ K}^{-1}$ To find the Gibbs free energy of formation at room temperature, recall that $G^0 = H^0 - T S^0$
 $= 92 \text{ kJ mol}^{-1} + (298 \text{ K}) (0.192 \text{ kJ mol}^{-1} \text{ K}^{-1}) = 35 \text{ kJ mol}^{-1}$
Alternatively, one can also find the temperature for which $G = 0$, $T = H_f^0 / S_f^0$.

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1 Basic Problems of Chemical Reaction Engineering and ... example problems before going on to other solutions. P2-5. This is a reasonably challenging problem that reinforces Levenspiel's plots. P2-6. Straight forward problem

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