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measurement has two significant figures. The measurement has four significant figures. Decimal point is present, so these zeros are significant. The rules are summarized in the following flowchart: qq Nonzero digits 3 4 0 0 g of sulfur pp Nonzero digits 4 0 0 0. mL of oxygen pp

CHEMFILE MINI-GUIDE TO PROBLEM SOLVING CHAPTER 2 ...

Name Date Class 1 of 13. CHEMFILE MINI-GUIDE TO PROBLEM SOLVING CHAPTER 8. Stoichiometry. So far in your chemistry course, you have learned that chemists count quantities of elements and compounds in terms of moles and that they re- late moles of a

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substance to mass by using the molar mass. In addition, you have learned to write chemical equations so that they represent the rearrangements of atoms that take place during chemical reactions, and you have learned to balance these equations.

CHEMFILE MINI-GUIDE TO PROBLEM SOLVING CHAPTER 8 Stoichiometry

Name Date Class 2 of 8 CHEMFILE MINI-GUIDE TO PROBLEM SOLVING SAMPLE PROBLEM 1 What is the molarity of a solution that is made by diluting 50.00 mL of a 4.74 M ...

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CHAPTER 15 ...

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Determine the percentage composition of each of the following compounds: a. sodium oxalate, $\text{Na}_2\text{C}_2\text{O}_4$ ans: 34.31% Na, 17.93% C, 47.76% O b. ethanol, $\text{C}_2\text{H}_5\text{OH}$ ans: 52.13% C, 13.15% H, 34.72% O c. aluminum oxide, Al_2O_3 ans: 52.92% Al, 47.08% O d. potassium sulfate, K_2SO_4 ans: 44.87% K, 18.40% S, 36.72% O

2.

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CHAPTER 6 ...

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General Plan for Solving Molarity Problems Mass of

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solute in g 1 Amount of solute in mol M moles solute liter solution 2 Volume of solution in L 3 Molar concentration, M 4 Convert using the molar mass of the solute. MOLARITY Molarity is the most common way to express concentration in chemistry.

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CHAPTER 14 ...

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Neither pressure nor temperature is constant; therefore, the combined gas law must be used to calculate the volume of CO₂ at the final temperature and pressure. multiply by the volume ratio, CO₂ use the combined gas law to convert from the initial

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temperature and pressure to the final temperature and pressure C 2 H 4

CHEMFILE MINI-GUIDE TO PROBLEM SOLVING
CHAPTER 13 ...

CHEMFILE MINI-GUIDE TO PROBLEM SOLVING
CHAPTER 14 ... CHEMFILE MINI-GUIDE TO
PROBLEM SOLVING Neither pressure nor temperature
is constant; therefore, the combined gas law must be
used to calculate the volume of CO₂ at the final tem-
perature and pressure. multiply by the volume ratio, CO
2 use the combined gas law to convert from the

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CHAPTER Empirical Formulas ADDITIONAL

PROBLEMS PRACTICE. Name Date Class 10 of 11

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Determine the empirical formula for compounds that

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CHEMFILE MINI-GUIDE TO PROBLEM SOLVING 1.

Carry out the following computations, and express the result in scientific notation: a. $2.290 \times 10^7 \text{ cm} \times 4.33 \times 10^3 \text{ s}$ ans: $5.29 \times 10^3 \text{ cm/s}$ b. $1.788 \times 10^5 \text{ L} \times 7.111 \times 10^3 \text{ m}^2$ ans: $2.514 \times 10^3 \text{ L/m}^2$ c. $5.515 \times 10^4 \text{ L} \times 6.04 \times 10^3 \text{ km}$ ans: 9.13 L/km d. $3.29 \times 10^4 \text{ km} \times 1.48 \times 10^2 \text{ min}$ ans: $2.22 \times 10^2 \text{ km/min}$

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I Chemfile Mini Guide To Problem Solving Ideal Gas Law Answers know that it is a time consuming job to

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Chemfile Mini Guide To Problem Solving Ideal Gas Law Answers

Name Date Class 4 of 13 CHEMFILE MINI-GUIDE TO PROBLEM SOLVING 3. Aluminum will react with sulfuric acid in the following reaction. $2\text{Al}(s) + 3\text{H}_2\text{SO}_4(l) \rightarrow \text{Al}_2(\text{SO}_4)_3(aq) + 3\text{H}_2(g)$ a. How many moles of H_2SO_4 will react ans: 27 mol H_2SO_4 with 18 mol Al? b. How many moles of each product ans: 27 mol H

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will be produced? 2 9 mol Al 2 (SO 4) 3 4.

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