

## Chem Fax Mole Ratios Answers

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How to Do Solution Stoichiometry Using Molarity as a Conversion Factor | How to Pass Chemistry *What You Need to Know to Pass a Test on Stoichiometry, Mole to Mole Ratios, and Avogadro's Number Calculating Moles in a Balanced Equation with the Mole Ratio* How to Find the Mole Ratio and Molar Mass *Construction of Mole Ratio as Conversion Factor- First Year Chemistry*

*Chapter No.1 Chemistry Lesson: Mole Ratio from Chemical Formulas* CHEM 2- Chapter 4 Moles Ratios in Chemical Formulas

Stoichiometry-Mole Ratios.mov *Mole ratio -Intro to stoichiometry* Chem Fax Mole Ratios Answers

First we determine the mole ratio (remember unknown species on top), H 2 S: H 2 = 8:8 = Number of moles of given species x mole ratio = number of moles required of unknown. So 3.5 x 1 = 3.5 moles of H 2 S could be produced. II. What would be the mass of this maximum yield? To answer this we need to convert our mole value for H 2 S into a mass value.

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moles Cu reacted = 0.1/63.5 = 1.574 x 10-3. 2. mass of silver = 52.1 - 50.1 = 2.00g. moles Ag produced = 2/108 = 8.19 x 10-1. 3. ratio Ag:Cu = 8.19 x 10-1 : 1.574 x 10-3 = 819 x 10-3 : 1.574 x 10-3...

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.... ..moles Cu reacted = 0.20g x (1 mol / 63.55g) = 3.15x10^-3 (2).. mass Ag produced = 50.67g - 50.03g = 0.64g.. ... ..mole Ag produced = 0.64g x (1 mol / 107.9g) = 5.93x10^-3 (3) dividing.. 5.93...

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