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- No two electrons in an atom can have the same unique set of four quantum numbers.
- The first quantum number describes

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the principal energy level. For example, the quantum number 2 identifies the second principal energy level. • The first two quantum numbers together describe a sublevel.

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by. ... it helped lay the foundation for the modern quantum theory.

- d.) it helps to locate a proton in an atom.
- d.) it is the same as the Bohr's theory ...
- the atomic nucleus
- b.) atomic orbitals
- c.) atoms
- d.) molecules.

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quantum theory.  
d.) it helps to locate a proton in an atom. d.) it helps to locate a proton in an atom.  
... the atomic nucleus b.) atomic orbitals c.) atoms d.) molecules.

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combustion.

energy. work.

kinetic energy. the

process of burning

something that

creates energy. the

capacity to do

work. what is done

to move an object

against some sort

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of resistance.... the  
energy of motion  
( $\frac{1}{2} mv^2$ )  
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three statements  
that make up the



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4 Modern atomic theory. 2. Define atomic number.

What is the atomic number for a boron atom? 3. Which elements are alkali metals, alkaline earth metals, and halogens? 4. Two of the most well-known isotopes of carbon are carbon-12 and

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carbon-14.

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Electron 417 the  
electron are also  
called orbitals. The  
orbital shown in  
Figure 11.3 is  
called the 1s  
orbital. The negative  
charge is most  
intense at the  
nucleus and  
decreases in  
intensity with  
distance outward.

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Theory (pages  
113-118). Use this  
table to answer  
questions 12 and  
13.. Name Chapter  
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the following  
questions, circle  
the best answer.  
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The modern atomic theory states that atoms of one element are the same, while atoms of different

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elements are different. What makes atoms of different elements different? The fundamental characteristic that all atoms of the same element share is the number of protons. All atoms of hydrogen have one and only one

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proton in the  
nucleus; all atoms  
of iron have 26  
protons in the  
nucleus.

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Important  
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separate atoms  
(Figure 4.1).

Radiant energy  
from the sun  
causes sunburn 4.1

Energy Some  
reactions absorb  
energy. Energy  
from the sun  $O_2(g)$



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+ energy 20(g)

Figure 4.1 Some reactions absorb energy. 120

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Before we can begin to explain the role that energy plays in these and other chemical

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