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1 . 5.3 × × 10 3 C 3 . (a) reduction; (b) oxidation; (c) oxidation; (d) reduction 5 . (a) F 2 + Ca 2F □ + Ca 2+ ; F 2 + Ca 2F □ + Ca 2+ ; (b) Considerations include: cost of the materials used in the battery, toxicity of the various components (what constitutes proper disposal), should it be a primary or secondary battery, energy requirements (the □size□ of the battery/how long ...

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Rhonda_FrazierTEACHER. Glencoe Chemistry Chapter 17. reversible recation. law of chemical equilibrium. equilibrium constant expression. Keq> 1. chemical reaction that can occur in borh the forward and the r□. particular ratio of reactant and product concentrations has a□. keq= [C]^c [D]^d/ [A]^a [B]^b.

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17-7 K and the extent of reaction K reflects a particular ratio of product concentrations to reactant concentrations for a reaction. A small value for K indicates that the reaction yields little product before reaching equilibrium. The reaction favors the reactants. K therefore indicates the extent of a reaction, i.e., how far a reaction proceeds towards the products at a given

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17.6 Corrosion □ Chemistry
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1 Terms. Key Concepts: Terms in this set (25) Matter. Anything that has mass and takes up space. Chemistry. The study of the properties of matter and how matter changes. Substance.

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Chemistry, 11e (Brown/LeMay/Brusten/Murphy) Chapter 17: Additional Aspects of Aqueous Equilibria 8) Calculate the pH of a solution prepared by dissolving 0.250 mol of benzoic acid 7 5 2 (C H O H) and 0.150 mol of sodium benzoate 7 5 2 (NaC H O ) in water sufficient to yield 1.00 L of solution. The K a of benzoic acid is 5 6.50 10-×.