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Acid + base ? salt + water + heat. $H_2SO_4 + MgO \rightarrow MgSO_4 + H_2O$
 $2HCl + Mg(OH)_2 \rightarrow MgCl_2 + 2H_2O$.
2. Reaction of non-metal oxides with bases. Non-metal oxides are acidic in nature
Base + Non-metal oxide ? salt + water + heat.
 $2NaOH + CO_2 \rightarrow Na_2CO_3 + H_2O$. To know more about Properties of Acids and Bases, visit here. Water Acids and bases in water

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Class 10 Chapter 2 Science Notes Acids, Bases, and Salts

The reaction between an acid and a base to give a salt and water is known as a neutralisation reaction. In general, a neutralisation reaction can be written as -
Base + Acid → Salt + Water

2.1.5 Reaction of Metallic Oxides with Acids
Activity 2.7 n

Take a small amount of copper oxide in a beaker and add dilute hydrochloric acid slowly while stirring.

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CHAPTER2 Acids, Bases and Salts

CBSE Class 10 Science Notes Chapter 2 Acids Bases and Salts. 1. Common Salt (Sodium Chloride): Sodium chloride (NaCl) is also known as Common or Table Salt. It is formed after the reaction between sodium ... 2. Bleaching Powder (CaOCl_2): Bleaching powder is also known as chloride of lime. It is a ...

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Notes Chapter - 2 Acids, Bases and Salts

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ACIDS: • These are the substances which have sour in taste. • They turns blue litmus solution to red. • They give H ions in aqueous solution Examples of Acids : - HCl - Hydrochloric Acid BASES • These substances are bitter in taste. • They turns red litmus solution to blue.

Chapter - 2, Acids Bases and Salts Notes
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Bases and Salts Acids Bases and Salts Class
10 Notes Understanding the Lesson 1. Acids
and Bases: Acids are sour in taste and change

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the colour of blue litmus to red.

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Chapter 2

Salts . When acid and base neutralize, salts are formed. Strong acid and strong base combines to form neutral salt. $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$. Eq.1. Formation of Neutral Salt. Strong acid and weak base combine to form acidic salt. For Example, Hydrochloric Acid and ammonium hydroxide combine to form ammonium chloride. Other examples, sodium hydrogen carbonate, sodium hydrogen sulphate

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etc.

Revision Notes for Science Chapter 2 - Acids, bases and ...

Study Material and Notes of Ch 2 Acids, Bases and Salts Class 10th Science. ? The sour and bitter tastes of food are due to acids and bases are present in them. ? Acids are sour in taste and change the colour of blue litmus to red. ? Litmus solution is a purple dye, which is extracted from lichen.

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Chemicals from common salt; NCERT Solutions for Class 10 Chapter 2 Science Acids bases and salts. In this chapter, students get hold of basic knowledge on Acids bases and salts. In this chapter, various chemical properties of acids, bases and salts and the reaction of them with metals, non-metals are studied.

NCERT Solutions for class 10 chapter 2 Acids, bases and salt

When an acid and base react with each other

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to give a salt and water, it is termed as neutralisation reaction i.e., Acid + Base ? Salt + Water For example: $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$. $\text{HNO}_3 + \text{KOH} \rightarrow \text{KNO}_3 + \text{H}_2\text{O}$. Question 15. Give two important uses of washing soda and baking soda. Answer: Important uses of washing soda are:

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And Bases has been provided by India's topmost Chemistry teachers. Also in this article, you will find the Step-wise explanation for each and every question.

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(ii) Acids react with metal carbonate and metals hydrogen carbonate to give salt, water and carbon dioxide. (iii) Acids react with bases to give salt and water. This reaction is called neutralization reaction. (iv) Acids react with metals oxides to give salt and

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water.

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Class 10 Chemistry Chapter 2 Important Questions with Answers Acids, Bases and Salts Acids, Bases and Salts Class 10 Important Questions Very Short Answer Type Question 1. How is the concentration of hydronium (H_3O^+) ions affected when a solution of an acid is diluted? (2017 OD)

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Salts dissociate in water fully and form
their basic and acidic counterpart. For

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example here CH_3COONa will form CH_3COOH and NaOH with water. Since NaOH is highly basic, the solution will be basic. N.b. the reverse of this also applies. For example, a solution of a salt of a strong acid and weak base will be acidic.

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Acids Bases and Salts Extra Questions Very Short Answer Type. Question 1. Two solutions A and B have pH values of 5 and 8 respectively. Which solution will be basic in nature ...

Acids Bases and Salts Class 10 Extra Questions with ...

For example, Hydrochloric acid (HCl), Sulphuric acid (H_2SO_4), Nitric acid (HNO_3). Bases - Bases are compounds which give hydroxide ion in water solution. For example, Sodium hydroxide (NaOH), Potassium hydroxide

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(KOH), Calcium hydroxide ($\text{Ca}(\text{OH})_2$). Salts - Salts are compounds obtained by replacement of H from acid by a metal.

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